\*Elements are pure substances which are made of one type of atom.

\*Atoms are neutral, which means they have the same amount of protons as they have electrons.

[An example: Helium (2nd on the periodic table) means it has 2 protons and 2 electrons.]

\*An ion of a charged atom; the protons are not equal to the number of electrions.

\*Atomic number refers to the amount of protons.

\*Mass number refers to the amount of protons and electrons.

\*The ionisation number decreases as you go down due to the atomic radius gets further from the nucleus.

\*The ionisation number increases as you go across due to the number of positive protons increasing.

Electronegativity

\*Electronegativity is the measure of attraction for a paired share of electrons.

\*If a difference between two elements is bigger than 1.7, an ion is formed (Ionic Bonding)

\*Polar covalent means “Will it be water soluble?”. This means it will dissolve in water.

\*Things that have a high binding point will also have a high melting point. Like salt.

\*Isotope refers to having the same number of protons as well as a different number of neutrons.

\*Relative atomic mass pertains to the average mass of an atom.